

Unit 7 Chemical Reactions Test Review

product

Part I: Conservation of Mass



1. If you react 5.2 grams of magnesium metal with 2.3 grams of copper (II) bromide, how many total grams of product will you produce?

$$5.2 + 2.3 = 7.5 \text{ g product}$$

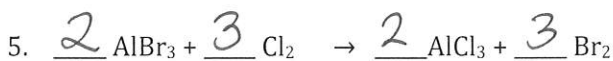
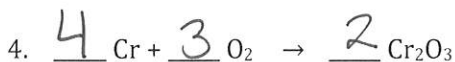
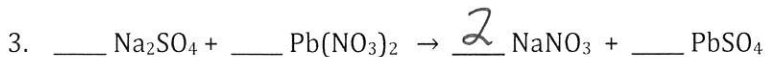
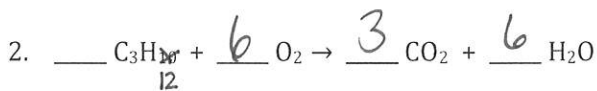
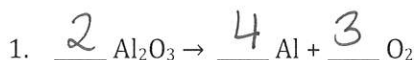
2. If 4.9 grams of magnesium metal react with copper (II) bromide to produce 6.2 grams of magnesium bromide and 3.3 grams of copper metal, how many grams of copper (II) bromide were used up?

$$4.9 + ? = 6.2 + 3.3$$

$$4.9 + ? = 9.5$$

$$4.6 \text{ g CuBr}_2$$

Part II: Balance the following reactions, identify what type of reaction, and write if the reaction is considered a redox reaction.

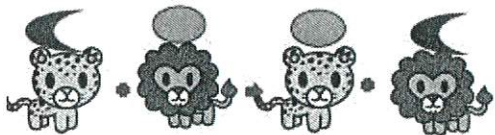


Type:	Redox?
1) D	✓
2) C	✓
3) DR	X
4) S	✓
5) SR	✓

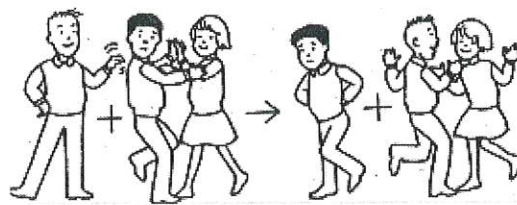
e⁻ transfer

Part III: Identify the type of reaction pictured.

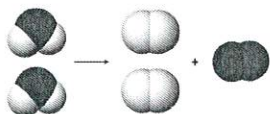
(pictures never have combustion)



Type of Reaction: DR



Type of Reaction: SR

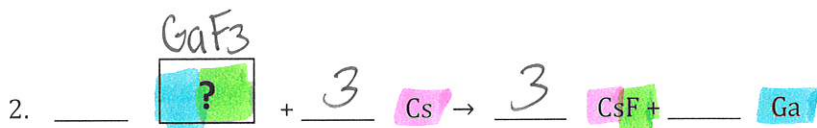
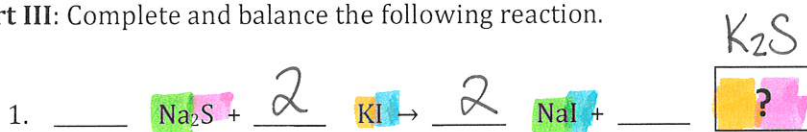


Type of Reaction: D



Type of Reaction: S

Part III: Complete and balance the following reaction.



Choose from the following options to identify the missing product in the reaction to the left.

1) NaK KSO₄ K₂S Si₂

2) CsGa F₂ GaF₃ CsF

Part IV: Oxidation-Reduction (Redox) Reactions

1. In an oxidation-reduction reaction, $\underline{e^-}$ are lost by one atom and gained by another atom.

2. If 4 electrons are lost in an oxidation-reduction reaction, how many are gained? 4e⁻

redox = transfer of e⁻

all types except DR